Structure Refinement of Mathiasite, $\left(\mathrm{K}_{0.62} \mathrm{Na}_{0.14} \mathrm{Ba}_{0.14} \mathrm{Sr}_{0.10}\right)_{\sum 1.0}\left[\mathrm{Ti}_{12.90} \mathrm{Cr}_{3.10} \mathrm{Mg}_{1.53} \mathrm{Fe}_{2.15} \mathrm{Zr}_{0.67} \mathrm{Ca}_{0.29}{ }^{-}\right.$ $\left.(\mathbf{V}, \mathrm{Nb}, \mathrm{Al})_{0.36}\right]_{\sum 21.0} \mathrm{O}_{38}$<br>By Bryan M. Gatehouse<br>Department of Chemistry, Monash University, Clayton, Victoria, Australia 3168<br>Ian E. Grey<br>CSIRO Division of Mineral Chemistry, PO Box 124, Port Melbourne, Victoria, A ustralia 3207<br>and Joseph R. Smyth<br>Geological Research Group, Los Alamos Scientific Laboratory, Los Alamos, New Mexico 87545, USA

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#### Abstract

M_{r}=1693.08\), rhombohedral, $R \overline{3}, a=$ 9.119 (3) $\AA, \alpha=69.24$ (2) $)^{\circ}, Z=1, D_{x}=4.39 \mathrm{~g} \mathrm{~cm}^{-3}$, $\lambda(\mathrm{Mo} K \alpha)=0.71069 \AA, \mu=68.8 \mathrm{~cm}^{-1}$. Final $R=$ 0.054 for 892 observed unique diffractometer data. Mathiasite is isostructural with the crichtonite-group minerals, $A M_{21} \mathrm{O}_{38}$, and is characterized by dominant potassium in the large-cation $A$ site. Partial disorder in the $M$-cation sublattice is interpreted as due to the partial occupation of a second anion site by large cations.


Introduction. New rhombohedral $\mathrm{Ba}-\mathrm{Sr}-\mathrm{K}-\mathrm{Cr}-\mathrm{Fe}$ titanate minerals, related to the crichtonite-group minerals (Gatehouse, Grey, Campbell \& Kelly, 1978), have been recently identified in peridotite nodules from the Bultfontein kimberlite in South Africa (Smyth, Erlank \& Rickard, 1978). The compositions conform to the general formula $A M_{21} \mathrm{O}_{38}, A=$ large cations (occupying a 12 -coordinate anion site in crichtonite) and $M=$ small cations. Two new members, with $A=$ dominant Ba and K have been provisionally named lindsleyite and mathiasite respectively. In this paper we report the structure refinement for the K -dominant member, mathiasite.

Electron microprobe analyses (average of three points) gave Ti, $36 \cdot 7$; $\mathrm{Cr}, 9 \cdot 58$; $\mathrm{Fe}, 7 \cdot 17$; $\mathrm{Zr}, 3 \cdot 65$; Mg , $2 \cdot 21$; K, 1.44; Ba, 1.16; Ca, 0.69; Sr, 0.52; Nb, 0.55; $\mathrm{V}, 0.39$; Al, $0.18, \mathrm{Na}, 0.18 \mathrm{wt} \%$. The corresponding unit-cell composition, normalized to 38 oxygens, and with the $\mathrm{Fe}^{2+} / \mathrm{Fe}^{3+}$ ratio adjusted to give 22 cations is: $\left(\mathrm{K}_{0 \cdot 62} \mathrm{Na}_{0 \cdot 14} \mathrm{Ba}_{0 \cdot 14} \mathrm{Sr}_{0 \cdot 10}\right)\left[\mathrm{Ti}_{12 \cdot 90} \mathrm{Cr}_{3 \cdot 10} \mathrm{Mg}_{1 \cdot 53} \mathrm{Fe}_{2} \cdot{ }_{15} \mathrm{Zr}_{0 \cdot 67}\right.$ $\mathrm{Ca}_{0 \cdot 29}\left(\mathrm{~V}, \mathrm{Nb}, \mathrm{Al}_{0} \cdot \cdot 36 \mathrm{O}_{38}\right.$ i.e. $A M_{21} \mathrm{O}_{38}$ with $A=$ large cations, K, Na, Ba, Sr.

Experimental. Crystal fragment measuring $0.044 \times$ 0108-2701/83/040421-02\$01.50
$0.069 \times 0.125 \mathrm{~mm}$ used for intensity data collection, PW1100 automated four-circle diffractometer, graphite-monochromated Mo $K \alpha, \omega-2 \theta$ scans, $\theta$ range $3-30^{\circ}$, scan speed $0.05^{\circ} \mathrm{s}^{-1}$, variable scan width $\Delta \theta=(1.20+0.3 \tan \theta)^{\circ}$; lattice parameters obtained from average of 15 orientation matrices from data collection; data corrected for absorption, maximum and minimum transmission factors 0.705 and $0.323 ; 5027$ $( \pm h, k, l)$ reflections collected and reduced to 1251 unique reflections, of which 892 with $I>3 \sigma(I)$ used in the refinement; atomic coordinates for loveringite (Gatehouse et al., 1978) in space group $R \overline{3}$ used as starting parameters, and the cations ordered into sites $M(0)-M(5)$ in order of decreasing size, as found for other crichtonite-group minerals; full-matrix leastsquares refinement of the scale factor, positional and isotropic thermal parameters converged at $R=0.06$.

The temperature factor for $O$ (7) was almost zero and a difference Fourier map revealed a subsidiary peak at $0.37,0.37,0.37$, of height $3 \mathrm{e} \AA^{-3}$. Bond-length calculations confirmed that this corresponded to an octahedrally coordinated atom, $\overline{M-O}=2 \cdot 12 \AA$, which face-shared to the $M(2)$ tetrahedron on the trigonal axis. This site, $M(6)$, was allowed partial occupancy by Mg and a corresponding reduction in the site occupancy of $M(2)$ was introduced. The population parameters for $M(6)$ and $O(7)$ were refined and finally, a refinement of all coordinates and isotropic thermal parameters was carried out, resulting in convergence at $R=0.054, R_{w}=0.045\left[w=1 / \sigma^{2}(F)\right]$ for the observed reflections.

Scattering factors for neutral atoms were those of Cromer \& Mann (1968), anomalous-dispersion corrections for all atoms from Cromer \& Liberman (1970); SHELX 76 (Sheldrick, 1976) and XRAY 76 (Stewart, 1976) programs were used for all computing.
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Discussion. Final positional and isotropic temperature factors are given in Table 1.* Bond lengths associated with the different polyhedra are given in Table 2.

* Lists of structure factors have been deposited with the British Library Lending Division as Supplementary Publication No. SUP 38283 ( 6 pp.). Copies may be obtained through The Executive Secretary, International Union of Crystallography, 5 Abbey Square, Chester CH1 2HU, England.

Table 1. Fractional coordinates and isotropic thermal parameters

|  | Occupancy | $x$ | $y$ | $z$ | $B\left(\dot{\AA}^{2}\right)$ |
| :---: | :---: | :---: | :---: | :---: | :---: |
| $M(0)$ | $\begin{aligned} & 0.62 \mathrm{~K}+0.14 \mathrm{Na}+ \\ & 0.14 \mathrm{Ba}+0.10 \mathrm{Sr} \end{aligned}$ | 0 | 0 | 0 | 1.24 (6) |
| $M(1)$ | $0.7 \mathrm{Zr}+0.3 \mathrm{Ca}$ | $\frac{1}{4}$ |  | $\frac{1}{2}$ | 0.26 (3) |
| $M(2)$ | $0.9 \mathrm{Mg}+0.9 \mathrm{Fe}$ | 0.3111 (1) | 0.3111 (1) | 0.3111 (1) | 0.36 (5) |
| $M(3)$ | $\begin{aligned} & 0.9 \mathrm{Ti}+3.1 \mathrm{Cr}+ \\ & 1.3 \mathrm{Fe}+0.7(\mathrm{~V}, \mathrm{Nb}, \mathrm{Mg}) \end{aligned}$ | 0.3479 (2) | $0 \cdot 1227$ (2) | 0.0232 (2) | 0.58 (2) |
| $M(4)$ | 6 Ti | 0.3071 (2) | 0.7208 (2) | 0.1457 (2) | 0.38 (2) |
| $M(5)$ | 6 Ti | 0.4754 (2) | 0.0832 (2) | 0.6386 (2) | 0.38 (2) |
| $M(6)$ | 0.2 Mg | 0.3692 (26) | 0.3692 (26) | 0.3692 (26) | 2.23 (1.19) |
| $\mathrm{O}(1)$ |  | 0.3084 (7) | 0.6261 (7) | 0.3820 (7) | 0.52 (9) |
| $\mathrm{O}(2)$ |  | 0.1553 (7) | 0.2394 (7) | 0.9392 (6) | 0.52 (9) |
| $\mathrm{O}(3)$ |  | 0.9218 (7) | 0.4583 (7) | 0.2979 (7) | 0.52 (8) |
| $\mathrm{O}(4)$ |  | 0.1424 (7) | 0.5175 (7) | 0.9902 (7) | 0.52 (8) |
| $\mathrm{O}(5)$ |  | 0.3901 (7) | 0.4879 (7) | 0.1343 (7) | 0.37 (8) |
| O(6) |  | 0.7043 (6) | 0.2434 (7) | 0.0743 (7) | 0.42 (8) |
| $\mathrm{O}(7)$ | $P P=1.056$ (18) | 0.2133 (3) | 0.2133 (3) | 0.2133 (3) | 0.45 (14) |

Table 2. $M-\mathrm{O}$ bond lengths $(\AA)$

| $M(0)$ cuboctahedron |  |  | $M(4)$ octahedron (titanium) |  |
| :--- | :--- | :--- | :--- | :---: |
| $M(0)-\mathrm{O}(2) \times 6$ | $2.803(7)$ | $M(4)-\mathrm{O}(2)$ | $1.889(6)$ |  |
| $M(0)-\mathrm{O}(6) \times 6$ | $2.861(5)$ | $M(4)-\mathrm{O}(6)$ | $1.949(7)$ |  |
| Mean | 2.832 | $M(4)-\mathrm{O}(3)$ | $1.967(6)$ |  |
|  |  | $M(4)-\mathrm{O}(5)$ | $2.014(6)$ |  |
| $M(1)$ octahedron |  | $M(4)-\mathrm{O}(1)$ | $2.020(6)$ |  |
| $M(1)-\mathrm{O}(1) \times 6$ | $2.136(6)$ | $M(4)-\mathrm{O}(6)^{\prime}$ | $2.028(6)$ |  |
|  |  |  | 1.978 |  |
|  |  | $M(5)$ octahedron (titanium) |  |  |
| $M(2)$ tetrahedron |  | $M(5)-\mathrm{O}(4)$ | $1.885(7)$ |  |
| $M(2)-\mathrm{O}(5) \times 3$ | $1.969(6)$ | $M(5)-\mathrm{O}(1)$ | $1.902(6)$ |  |
| $M(2)-\mathrm{O}(7)$ | $2.019(4)$ | $M(5)-\mathrm{O}(3)$ | $1.930(6)$ |  |
| Mean | 1.981 | $M(5)-\mathrm{O}(5)$ | $1.987(7)$ |  |
|  |  | $M(5)-\mathrm{O}(5)^{\prime}$ | $2.007(7)$ |  |
| $M(3)$ octahedron |  |  | $2.114(8)$ |  |
| $M(3)-\mathrm{O}(4)$ | $1.934(6)$ |  | 1.971 |  |
| $M(3)-\mathrm{O}(3)$ | $1.963(7)$ | $M(6)$ octahedron |  |  |
| $M(3)-\mathrm{O}(2)$ | $1.963(6)$ | $M(6)-\mathrm{O}(5) \times 3$ | $2.003(20)$ |  |
| $M(3)-\mathrm{O}(4)^{\prime}$ | $1.992(6)$ | $M(6)-\mathrm{O}(1)+3$ | $2.244(25)$ |  |
| $M(3)-\mathrm{O}(7)$ | $1.995(3)$ | $M e a n$ | 2.123 |  |
| $M(3)-\mathrm{O}(2)^{\prime}$ | $2.032(6)$ |  |  |  |
| Mean | 1.980 |  |  |  |

The results of the structure refinement confirm that mathiasite is isostructural with other crichtonite-group minerals. The structure has been described in detail in previous publications (Grey, Lloyd \& White, 1976; Gatehouse, Grey \& Kelly, 1979). The cation-ordering scheme proposed, Table 1 , is supported by valence-sum calculations, using the parameters of Brown \& Wu (1976) and by the reasonable thermal parameters. The partial occupancy of the octahedral site $M(6)$ has not been observed in the refinements of other crichtonitegroup minerals. This structural feature is probably related to partial occupancy of the anion site on the trigonal axis, $\mathrm{O}(7)$, by large cations. Refinement of the population parameter for $O(7)$ gave a value significantly greater than $1[1.056(18)]$, corresponding, for example, to $0.15 \mathrm{Na}+0.85 \mathrm{O} . \mathrm{O}(7)$ is the apical oxygen of the $M(2) \mathrm{O}_{4}$ tetrahedron, and we propose that when $\mathrm{O}(7)$ is occupied by a large cation, the cation in $M(2)$ can no longer occupy the tetrahedral site and moves into the adjacent octahedral site $M(6)$. The octahedrally coordinated cations $M(3)$ also bond to $O(7)$, and the occupation of $O(7)$ by a large cation would result in either vacancies in $M(3)$, or in a displacement of the $M(3)$ cations away from $O(7)$ to take up a lower coordination. No evidence of the latter was obtained from the difference Fourier map.

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# Tetradecapraseodymium Hexanickel Undecasilicide, $\mathbf{P r}_{14} \mathbf{N i}_{6} \mathbf{S i}_{11}$, with Centered Trigonal Rare-Earth Prisms 

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Abstract. \(\quad M_{r}=2634, \quad\) monoclinic, \(\quad C 2 / m, \quad a=113.72(1)^{\circ}, \quad V=2809.7 \AA^{3}, \quad Z=4, \quad D_{x}=\) \(33.991(3), \quad b=4.2328(4), \quad c=21.330(3) \AA, \quad \beta=6.226 \mathrm{Mg} \mathrm{m}^{-3}, \quad\) Mo \(K \alpha, \quad \lambda=0.71069 \AA, \quad \mu=\) 0108-2701/83/040422-04\$01.50 © 1983 International Union of Crystallography
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